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~~19.1 — Acid-Base Theories.~~ Acids have a sour taste, change the color of an indicator, can be strong or weak electrolytes in aqueous solution, and react with metals.
Bases.

~~Chapter 19 — Acids, Bases, and~~

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Chapter 19 “ Acids, Bases, and Salts ... Thus, a conjugate acid-base pair is related by the . loss or gain. of a single hydrogen ion. Chemical Indicators? They are weak acids or bases that have a different color from their original

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~~Problems~~
acid and base. Acids and bases
come in pairs.

~~Chapter 19 Acids, Bases, and Salts~~
Chapter 19: Acids, Bases, and
Salts. STUDY. PLAY. Tastes sour.
Acid. Changes the color of an acid-
base indicator (acid, base, or both)

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Both. Can be strong or weak electrolytes in aqueous solution.
Acid. In drinks, citrus, pop, etc.
Acid. This acid is associated with protein. amino acid. Used in fragrances and flavors.

~~Chapter 19: Acids, Bases, and~~

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44 terms. Taneya_Heath. Chapter 19-Acids, Bases, and Salts.

STUDY. PLAY. Characteristics of Acids. they taste sour, will change the color of an acid-base indicator, and can be strong or weak electrolytes in an aqueous solution. Characteristics of Bases.

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When these acids and bases are mixed in the right proportions, the

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neutralization reaction thus results in the formation of salt and water. Some naturally occurring salts found in nature include NaCl and KCl etc in seawater and natural rock deposits.

~~Acids, Bases, and Salts~~

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~~Introduction, Dissociation ...~~

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Acids: taste sour, react with
certain metals to produce
hydrogen gas, react with
carbonates/bicarbonates to

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Problems
produce carbon dioxide gas. Corrode metals, pH is less than 7, turns blue litmus paper red. Bases: taste bitter, have a slippery feeling. React with acids to form salts and water, pH is greater than 7.

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~~Chapter 19: Acids, Bases, & Salts~~ ~~—Weebly~~

Blue litmus paper turns red in contact with an acid. Acids React with Carbonates and Bicarbonates.
 $\text{HCl} + \text{NaHCO}_3 \rightarrow \text{NaCl} + \text{H}_2\text{O} + \text{CO}_2$. Hydrochloric acid + sodium bicarbonate. salt + water + carbon

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dioxide. An old-time home remedy
for relieving an upset stomach.
OPERATIONAL BASES.

~~Chapter 19 Acids, Bases, and Salts
—MOLEBUS (ALLCHEM)~~

Chapter 19 Acids, Bases, and
Salts. Acids and Bases: Basics.

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Problems

These are really just a specific type of chemical compound. No mysteries here. 1) They MUST be ionic compounds, and. MUST dissolve in water, that ' s the only way to be chemically active. We call this. disassociation.

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Acid + base → salt + water +

heat. $\text{H}_2\text{SO}_4 + \text{MgO} \rightarrow \text{MgSO}_4$

$+ \text{H}_2\text{O}$. $2\text{HCl} + \text{Mg}(\text{OH})_2 \rightarrow \text{MgCl}_2$

$+ 2\text{H}_2\text{O}$. 2. Reaction of non-

metal oxides with bases. Non-

metal oxides are acidic in nature.

Base + Non-metal oxide → salt +

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Problems
water + heat. $2\text{NaOH} + \text{CO}_2 \rightarrow \text{Na}_2\text{CO}_3 + \text{H}_2\text{O}$. To know more about Properties of Acids and Bases, visit here.

~~Class 10 Chapter 2 Science Notes
Acids, Bases, and Salts~~

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Neutralization Reactions >

Chapter 19 Acids, Bases, and Salts

19.1 Acid-Base Theories 19.2

Hydrogen Ions and Acidity 19.3

Strengths of Acids and Bases 19.4

~~19.4 (1).ppt - 19.4 Neutralization
Reactions > Chapter 19 ...~~

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Acids Neutralize Bases $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$ -Neutralization reactions ALWAYS produce a salt (which is an ionic compound) and water. -Of course, it takes the right proportion of acid and base to produce a neutral salt fSulfuric Acid = H_2SO_4

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~~Chapter 19 Acids_Bases_and
Salts | Acid | Ph | Free 30 ...~~

weak acid and one of its salts. or. a
. weak base and one of its salts.
Buffers. A buffer system is better
able to resist changes in pH than
pure water. Since it is a pair of

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~~Problems:~~ one chemical
neutralizes any . . . Chapter 19
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Potassium Chloride (KCl): It is formed after the reaction between potassium hydroxide (a strong base) and hydrochloric acid (a

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strong acid). (ii) Acidic Salts:
Salts which are formed after the reaction between a strong acid and weak base are called Acidic salts. The pH value of acidic salt is lower than 7.

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