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$\text{HOOC}-\text{COOH} \cdot x\text{H}_2\text{O}$, by titrating a known mass of the acid with a $0.100 \text{ mol dm}^{-3}$ solution of $\text{NaOH}(\text{aq})$. 0.795 g of ethanedioic acid was dissolved in distilled water and made up to a total ...

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[2] (b) Marit arranged the values she found in Table 12 into an energy cycle. $\text{C}_2\text{H}_4(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$ $\Delta H = -141 \text{ kJ mol}^{-1}$ $\text{C}_2\text{H}_4(\text{g}) + 3\text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{l})$ $\Delta H = -1560 \text{ kJ mol}^{-1}$ $\text{C}_2\text{H}_6(\text{g}) + 7/2\text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g}) + 3\text{H}_2\text{O}(\text{l})$ $\Delta H = -286 \text{ kJ mol}^{-1}$ (hydrogenation) (1)
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