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Substituting and multiplying

by the factor 10⁹ for the

density unit kg/m³, the

mass of the atmosphere is

determined to be $m =$

5.092×10^{18} kg Discussion

Performing the analysis with

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excel would yield exactly the same results. EES

Solution for final result:
 $a=1.2025166$
 $b=-0.10167$
 $c=0.0022375$
 $r=6377$
 $h=25$
 $m=4 \cdot \pi \cdot (a \cdot r^2 \cdot h + r \cdot (2 \cdot a + b \cdot r) \cdot h^2 / 2 + (a + 2 \cdot b \cdot r + c \cdot r^2) \cdot h^3 / 3 + (b + 2 \cdot c \cdot r) \cdot h^4 / 4 + c \cdot h^5 / 5) \cdot 1E+9$
1-7 Pressure,

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Barometer1-34C The pressure
relative to the atmospheric
pressure is called ...

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The results are: $(z) = a +$
 $bz + cz^2 = 1.20252 -$
 $0.101674z + 0.0022375z^2$
for the unit of kg/m^3 ,
(or, $(z) = (1.20252 -$
 $0.101674z + 0.0022375z^2)$)

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109 for the unit of kg/m^3
) where z is the vertical
distance from the earth
surface at sea level. At $z =$
7 km, the equation would
give $= 0.60 \text{ kg}/\text{m}^3$.

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$0.101674z + 0.0022375z^2$ for

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the unit of kg/m^3 , (or, ρ
(z) = $(1.20252 - 0.101674z +$
 $0.0022375z^2) \times 10^9$ for the
unit of kg/km^3) where z is
the vertical distance from
the earth surface at sea
level. At $z = 7$ km, the
equation would give $\rho = 0.60$

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